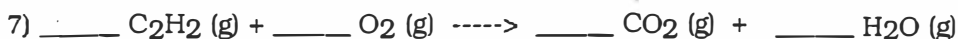
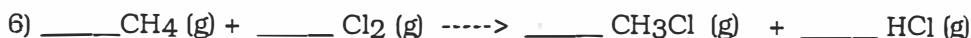
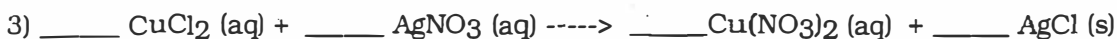
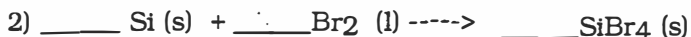
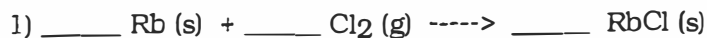


## Writing Balanced Equations

Name: \_\_\_\_\_

Part I: Balance the following equations by filling in the proper coefficients before the given formulas.

ALSO, IDENTIFY THE TYPE OF REACTION



Part II: Write balanced chemical equations for the following word reactions. ALSO, IDENTIFY THE TYPE OF REACTION

8) Oxygen gas plus nitrogen gas produces nitrogen dioxide (NO<sub>2</sub>) gas.

9) Aqueous beryllium iodide plus aqueous tin (IV) nitrate produces aqueous beryllium nitrate and solid tin (IV) iodide.

10) Solid iron plus fluorine gas yields solid iron (III) fluoride.

11) Aqueous hydrochloric acid (HCl) and aqueous magnesium hydroxide will mix and produce aqueous magnesium chloride and liquid water.

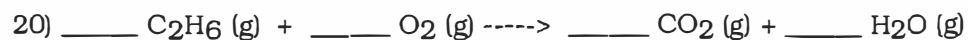
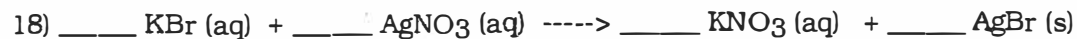
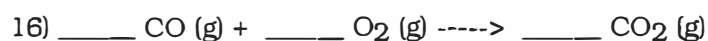
12) Aqueous ammonium nitride added to aqueous lead (II) chlorate reacts to give aqueous ammonium chlorate and solid lead (II) nitride.

13) Aqueous carbonic acid (H<sub>2</sub>CO<sub>3</sub>) decomposes to produce liquid water and carbon dioxide (CO<sub>2</sub>) gas.

14) Aqueous sodium phosphate plus aqueous copper (II) sulfate produces aqueous sodium sulfate and solid copper (II) phosphate

15) Liquid water combines with carbon monoxide (CO) gas to produce hydrogen gas and carbon dioxide (CO<sub>2</sub>) gas.

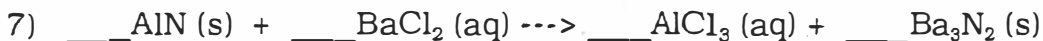
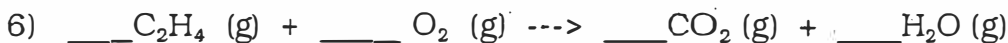
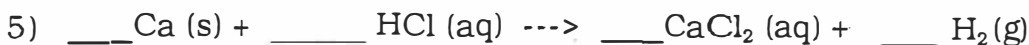
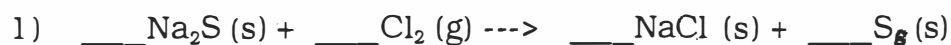
Part III: Balance the following equations. Then, write the chemical equation out in words.



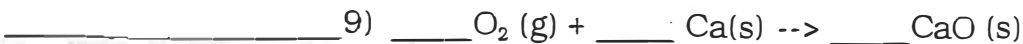
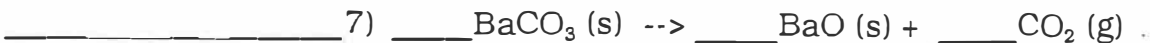
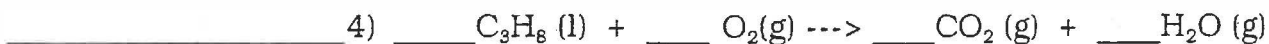
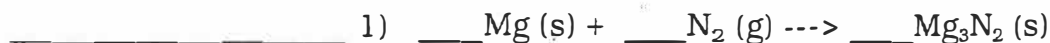
## Balancing Equations 2

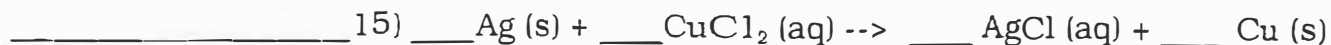
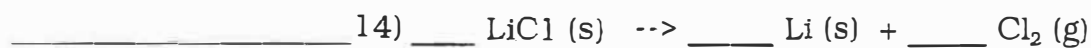
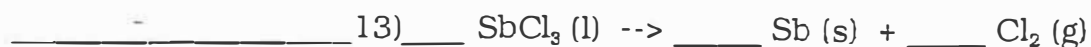
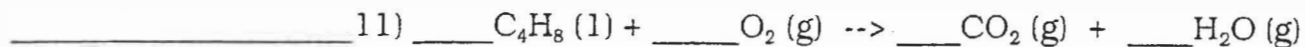
Name: \_\_\_\_\_

Part 1: Balance the following equations by filling in the proper coefficients before the given formulas.



Part 2: Balance the equations and write the type of reaction in the blank to the left.





Part 3: Balance the following equations. Then, write the chemical equation out in words.

