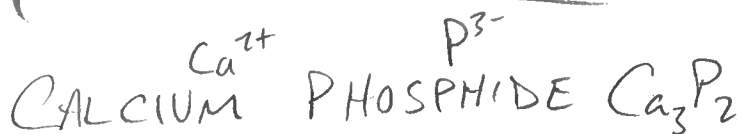


Th 2025-03-06 X&Y

TRY THESE

Tu 2024-03-12  
CHEM-3  
Y3/12

NAMES → FORMULAS



ALUMINUM SULFIDE

BARIUM CHLORIDE

4 POTASSIUM NITRIDE

IRON (II) IODIDE

IRON (III) CHLORIDE

TIN (II) OXIDE

TIN (IV) OXIDE

COBALT (II) NITRIDE

10 COBALT (III) SULFIDE

POTASSIUM CHLORATE

MAGNESIUM CHLORITE  $Mg(ClO_2)_2$

ALUMINUM PHOSPHATE  $AlPO_4$

CALCIUM PHOSPHATE

SODIUM PHOSPHATE

16 IRON (III) CARBONATE

FORMULAS → NAMES

$CaS$  CALCIUM SULFIDE

$B_2O_3$

$SrBr_2$

20  $Li_2Se$

$Cu_2S$

$CuCl_2$

$NiO$

$Ni_2O_3$

$CrCl_3$

26  $CrO$

$NaC_2H_3O_2$

$K_2C_2O_4$

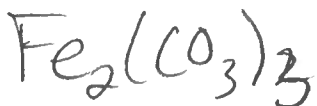
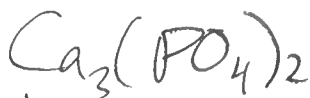
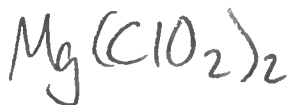
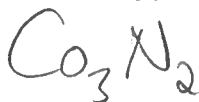
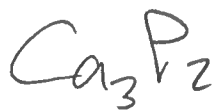
$CaSO_4$

$Mg(HCO_3)_2$

$Al_2(CO_3)_3$

32  $(NH_4)_3PO_4$

# ANSWERS



CALCIUM SULFIDE

BORON OXIDE

STRONTIUM BROMIDE

LITHIUM SELENIDE

COPPER (I) SULFIDE

COPPER (II) CHLORIDE

NICKEL (II) OXIDE

NICKEL (III) OXIDE

CHROMIUM (III) CHLORIDE

CHROMIUM (II) OXIDE

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SODIUM ACETATE

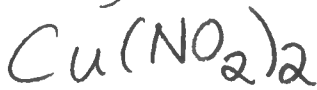
POTASSIUM OXALATE

CALCIUM SULFATE

MAGNESIUM HYDROGEN  
CARBONATE

ALUMINUM CARBONATE

AMMONIUM PHOSPHATE

PRACTICEFORMULASNAMES

COPPER (I) CARBONATE

IRON (II) CHROMATE

IRON (III) OXALATE

POTASSIUM CHLORITE

MAGNESIUM CHLORATE

 $\text{Na}_3\text{PO}_4$  SODIUM PHOSPHATE $\text{Ca}_3(\text{PO}_3)_2$  CALCIUM PHOSPHITE $\text{Ba}_3\text{P}_2$  BARIUM PHOSPHIDE $\text{Cu}(\text{NO}_2)_2$  COPPER (II) NITRITE $\text{Al}(\text{ClO})_3$  ALUMINUM HYPOCHLORITE $\text{Cu}_2\text{CO}_3$  COPPER (I) CARBONATE $\text{FeCrO}_4$  IRON (II) CHROMATE $\text{Fe}_2(\text{C}_2\text{O}_4)_3$  IRON (III) OXALATE $\text{KClO}_2$  POTASSIUM CHLORITE $\text{Mg}(\text{ClO}_3)_2$  MAGNESIUM CHLORATE

## PRACTICE WRITING NAMES &amp; FORMULAS

X &amp; Y

## MOLECULAR COMPOUNDS

<u>NAME</u>	<u>FORMULA</u>
1 SELENIUM TETRAFLUORIDE →	
2 ←	XeO <sub>4</sub>
3 ARSENIC TRIBROMIDE →	
4 ←	ClF <sub>5</sub>
5 DICHLORINE MONOXIDE →	
6 ←	N <sub>2</sub> O <sub>5</sub>
7 SULFUR HEXAFLUORIDE →	
8 ←	P <sub>4</sub> O <sub>10</sub>

DIDN'T FINISH IN GROUP X  
M2025-03-10

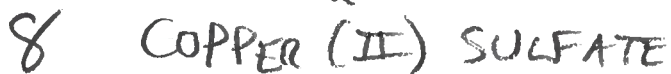
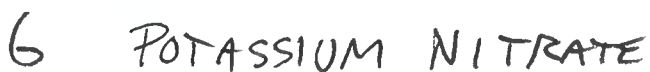
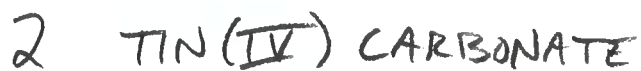
## MIX AND MATCH: IONIC AND MOLECULAR COMPOUNDS

<u>NAME</u>	<u>FORMULA</u>	<u>NAME</u>	<u>FORMULA</u>
1 ZINC CHLORIDE →		5 KRYPTON DIFLUORIDE →	
2 ←	Sn(CO <sub>3</sub> ) <sub>2</sub>	6 ←	KNO <sub>3</sub>
3 MAGNESIUM NITRIDE →		7 IRON(III) NITRITE →	
4 ←	SF <sub>4</sub>	8 ←	CuSO <sub>4</sub>

## ANSWERS



~~XXXXXXXXXX~~



W 2025-03-12 CHEM-3  
X & Y

# PRACTICE: NAMING MOLECULAR & IONIC COMPOUNDS

NAME	FORMULA	NAME	FORMULA
1 CALCIUM BROMIDE →		5 DIPHOSPHORUS PENTOXIDE →	
2 ← $Fe_3(PO_4)_2$		6 ← $Mg(NO_2)_2$	
3 POTASSIUM OXIDE →		7 LEAD(IV) NITRATE →	
4 ← $N_2O_4$		8 ← $Co_2(SO_3)_3$	
		9 $CF_4$ ↔ $CB_4$	

ANSWERS:

- |                        |                       |
|------------------------|-----------------------|
| 1 $CaBr_2$             | 5 $P_2O_5$            |
| 2 IRON(II) PHOSPHATE   | 6 MAGNESIUM NITRITE   |
| 3 $K_2O$               | 7 $Pb(NO_3)_4$        |
| 4 DINITROGEN TETROXIDE | 8 COBALT(III) SULFITE |
|                        | 9 CARBON TETRABROMIDE |

Write formulas for the following ionic compounds AND ACIDS

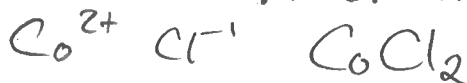
1. cobalt(II) chloride
2. cobalt(III) cyanide
3. copper(I) carbonate
4. copper(II) carbonate
5. silver phosphate
6. lead(II) hypochlorite
7. lead(IV) chlorite
8. iron(II) chlorate
9. iron(III) perchlorate
10. tin(II) chromate
11. tin(IV) dichromate
12. sodium sulfate
13. ammonium chloride
14. potassium nitrate
15. aluminum nitrate
16. barium hydrogen carbonate
17. barium carbonate
18. cobalt(III) sulfite
19. titanium(IV) sulfate
20. bismuth(III) acetate

21. NITRIC ACID

22. HYDROFLUORIC ACID

For EX.:

1. COBALT (II) CHLORIDE



## ANSWERS

Write formulas for the following ionic compounds.

- |   |  |
|---|--|
| <p>1. cobalt(II) chloride <math>\text{CoCl}_2</math></p> <p>2. cobalt(III) cyanide <math>\text{Co}(\text{CN})_3</math></p> <p>3. copper(I) carbonate <math>\text{Cu}_2\text{CO}_3</math></p> <p>4. copper(II) carbonate <math>\text{CuCO}_3</math></p> <p>5. silver phosphate <math>\text{Ag}_3\text{PO}_4</math></p> <p>6. lead(II) hypochlorite <math>\text{Pb}(\text{ClO})_2</math></p> <p>7. lead(IV) chlorite <math>\text{Pb}(\text{ClO}_2)_4</math></p> <p>8. iron(II) chlorate <math>\text{Fe}(\text{ClO}_3)_2</math></p> <p>9. iron(III) perchlorate <math>\text{Fe}(\text{ClO}_4)_3</math></p> <p>10. tin(II) chromate <math>\text{SnCrO}_4</math></p> | <p>11. tin(IV) dichromate <math>\text{Sn}(\text{Cr}_2\text{O}_7)_2</math></p> <p>12. sodium sulfate <math>\text{Na}_2\text{SO}_4</math></p> <p>13. ammonium chloride <math>\text{NH}_4\text{Cl}</math></p> <p>14. potassium nitrate <math>\text{KNO}_3</math></p> <p>15. aluminum nitrate <math>\text{Al}(\text{NO}_3)_3</math></p> <p>16. barium hydrogen carbonate <math>\text{Ba}(\text{HCO}_3)_2</math></p> <p>17. barium carbonate <math>\text{BaCO}_3</math></p> <p>18. cobalt(III) sulfite <math>\text{Co}_2(\text{SO}_3)_3</math></p> <p>19. titanium(IV) sulfate <math>\text{Ti}(\text{SO}_4)_2</math></p> <p>20. bismuth(III) acetate <math>\text{Bi}(\text{C}_2\text{H}_3\text{O}_2)_3</math></p> |
|---|--|

ACETATE  $\text{C}_2\text{H}_3\text{O}_2^-$  or  $\text{CH}_3\text{COO}^-$

